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Works Both  
Ways

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Video Le Chatelier's  
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~~Cobalt Complex Ions  
Hardy-Weinberg  
Equilibrium [SK015]  
Exp 5: Chemical  
Equilibrium~~

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~~Piaget's Theory of  
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Experiment #13: The  
Equilibrium Constant.~~

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~~Exploring Equilibrium  
Osmosis and Water  
Potential (Updated)~~

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FeSCN<sub>2</sub>+ Equilibrium

- LeChatelier's

Principle Lab Part 1

CHEM113L:

Equilibrium Constant

Post-lab Analysis

Chemical

Equilibrium: Colorful

Demonstrations

Equilibrium Lab

ANSWERS: Cobalt

CoCl<sub>4</sub><sup>2-</sup> and

Cu(H<sub>2</sub>O)<sub>6</sub><sup>2+</sup> Le

Chatelier's Principle

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of Chemical  
Equilibrium - Basic  
Introduction  
Introduction to

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Writing The  
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~~Equations: Crash~~

~~Course Chemistry #29~~

Spectrophotometric  
Determination of an

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Minute Le Chatelier's  
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~~EXPERIMENT 5:~~

~~Chemical Equilibrium  
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~~ANSWERS: Chromate/  
Dichromate~~

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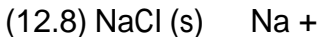
Experiment 5:  
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Keq FeSCN<sup>2+</sup> Lab  
Equilibrium Constant  
BCLN - Investigating  
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Lab Report Lab

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## Report On Exploring Equilibrium

In this lab, the effect of applying stresses to a variety of chemical systems at equilibrium will be explored. The equilibrium systems to be studied are given below:

Saturated Sodium  
Chloride Solution



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$1(aq) + Cl^{-} 1(aq)$

Equilibrium It  
12: Equilibrium and  
Le Chatelier's

Principle (Experiment

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Spectrophotometric



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Determination of  
Equilibrium Constant  
page 1 Lab 4 •

Spectrophotometric  
Determination of an  
Equilibrium Constant  
PURPOSE: To  
determine the value  
of the equilibrium  
constant for a  
reaction.

Lab Report On  
Exploring Equilibrium  
*Page 9/36*

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Go to the Reactions  
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## Lab Report On

### Exploring Equilibrium

#### It Works Both Ways

The position of the equilibrium for

reaction (1) is said to lie with the reactants, or to the left, because at equilibrium very little of the carbon dioxide has reacted.

On the other hand, in the reaction  $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$

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(2) the equilibrium position lies very far to the right since only very small amounts of  $H_2$  and  $O_2$

Laboratory 7:  
Chemical Equilibrium  
This preview shows  
page 1 - 2 out of 4  
pages. Exploring  
Equilibrium Mini-Lab  
Lab Write-up:  
Answers to

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Exploring Equilibrium It Works Both Ways

questions, graphs, data table. Group.

PART A. Login to the computer and open a web browser. Go to . Click “ Play with Sims, ” then find the “ Chemistry ” section, and choose “ Reactions and Rates. ” . When the simulation is open, click on the “ Many Collisions ” tab. Look

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at the screen and observe everything you can find out about the reaction pictured,  $A + BC \rightleftharpoons AB + C$ . 1) What type of ...

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## Exploring Equilibrium Lab #1

When red-colored  $\text{Co}(\text{NO}_3)_2 \times 6\text{H}_2\text{O}$  crystals are dissolved in 3mL of deionized water, the solution

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turns pink. To explain this, the equilibrium stress is on the product 's side (addition of water), so the solution shifts towards to reactants. Since the reactants are favored, the equilibrium produces more  $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$  ions, which results in a pink-colored

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 $\text{NaCl(s)} \rightleftharpoons \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$   
12:  
Equilibrium and Le

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Chatelier's Principle  
(Experiment...

Equilibrium It  
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Equilibrium is a condition in which all the forces in a system counteract each other. When adding vectors, equilibrium can be showed by a representative force, called an equilibrant.

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Exploring Equilibrium

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The average equilibrium constant was 474.76, because it is greater than one, at equilibrium, the reaction favor the formation of products. The method used to measure how much a chemical substance absorbs light by

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intensity of light as it  
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