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Two equilibria must be considered: CuO



+ Cu at equilibrium \rightleftharpoons



equilibrium \rightleftharpoons $n\text{O}_2$

$\rightleftharpoons 2$ In the final state,

only O_2 is present.

Therefore $\rightleftharpoons = n\text{O}_2$.

In an excess of O_2 , all the 1 copper is

oxidized. Therefore

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$\epsilon = 0$ or $\epsilon = n\pi$.

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