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(c) $\text{C-Cl} + \text{Cl-C}$

Chlorine is more

electronegative than

carbon Problem 1.6

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Draw Lewis
structures, showing all
valence electrons, for
these molecules.

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electronegativity $2.5 - 2.1 = 0.4$ $2.1 - 2.1 = 0$
 $4.0 - 2.5 = 1.5$ $3.0 - 2.5 = 0.5...$

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CHAPTER 1

Solutions to the

Problems Problem 1.1

Write and compare
the ground-state
electron

configurations for
each pair of elements:

(a) Carbon and silicon

C (6 electrons) $1s^2$

$2s^2 2p^2$ Si (14

electrons) $1s^2 2s^2$

$2p^6 3s^2 3p^2$ Both

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carbon and silicon
have four electrons in
their outermost
(valence) shells.

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...

William H. Brown is
emeritus professor of
chemistry at Beloit
College, where he
was twice named

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Teacher of the Year.

His teaching responsibilities include organic chemistry, advanced organic chemistry, and, more recently, special topics in pharmacology and drug synthesis.

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