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~~Calculating~~

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5.5-5.6,

5.10-5.11

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6.6-6.9 Enthalpy  
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oxidizing agent,

is bubbled into

a solution of

potassium

bromide.

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(a) Calculate the value of the standard free energy change,  $\Delta G_{298}$ , for the reaction. (b) Determine the temperature at

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Answers  
which the equilibrium constant,  $K_{eq}$ , for the reaction is equal to 1.00. (Assume that  $\Delta H_{rxn}$  and  $\Delta S_{rxn}$  are independent of temperature.)

(c) Calculate the standard enthalpy change,  $\Delta H_{rxn}$ , that occurs when a 0.256 mol

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B) Question 1

Sample: 1A

Score: 9 This  
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all 9 points: 1  
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for part (b), 2  
for part (c), 1

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for part (d), 2

for part (e),

and 1 for part

(f) Sample: 1B

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Sample: 1A

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Score: 6 The

point was not

earned in part

(a) because the

student

incorrectly

attempts to use

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