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product that forms when a reaction is carried out in the laboratory. excess reagent.

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The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

Stoichiometry is based on the law of conservation of mass.

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begin your rxn with 56.7g of zinc iodide?

56.7 g 1 mole 1 mole 253.8 g 319.19 g 1 mole

1 mole = 45.08 grams of Iodine 2. 1 $\text{Pb}_2(\text{CO}_3)_3$

+ 6 K \rightarrow 2 Pb + 3 K_2CO_3

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319.19 g 1 mole 1 mole = 45.08 grams of
Iodine 2. $1 \text{ Pb}_2(\text{CO}_3)_3 + 6 \text{ K} \rightarrow 2 \text{ Pb} + 3 \text{ K}_2\text{CO}_3$

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